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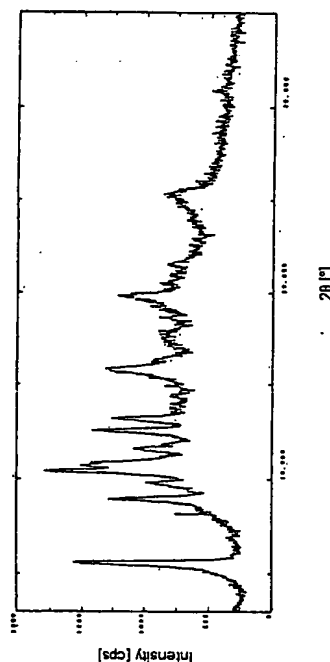
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(57) The present invention provides an E-type crystal of N-demethyl-N-isopropyl-12-methoxy-11-oxo-8,9-anhydroerythromycin A-6,9-hemiacetal fumarate having strong diffraction peaks at diffraction angles ( $2\theta$ ) of  $5.6^\circ$  and  $10.4^\circ$  as measured by powder X-ray diffraction, which is prepared by treating a C-type crystal of the compound in a mixed solvent of ethyl acetate and water at  $20^\circ\text{C}$  to  $40^\circ\text{C}$ , and a D-type crystal prepared via the E-type crystal. These crystals have a reduced content of residual solvent and high suitability for formulation.

Fig. 4



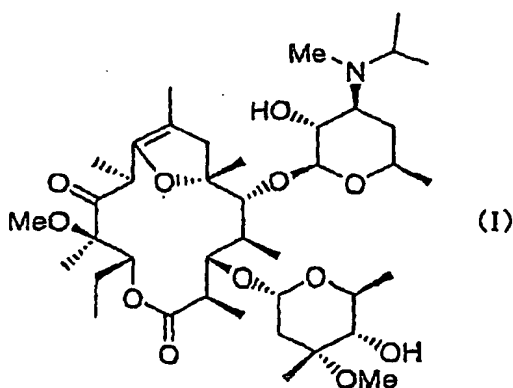
## Description

## TECHNICAL FIELD

[0001] The present invention relates to novel crystals of fumarate salts of erythromycin derivatives and a method for their preparation.

## BACKGROUND ART

[0002] The compound represented by formula (I):



is disclosed in, for example, JP 6-56873 A (WO93/24509) and JP 9-100291 A (WO97/06177). This compound is known to have the ability to enhance the movement of the digestive tract.

[0003] The preparation of this compound is disclosed in, for example, JP 9-100291 A, Bioorg. & Med. Chem. Lett. vol. 4(11), 1347 (1994) and JP 9-100291 A.

[0004] Conventionally, there are three patterns for crystals of a fumarate salt of compound (I): A-type crystal (hereinafter simply referred to as "A-type crystal"), C-type crystal (hereinafter simply referred to as "C-type crystal") and D-type crystal (hereinafter simply referred to as "D-type crystal"). Each of the A-type, C-type and D-type crystals is disclosed in JP 9-100291 A and can be prepared as described in this publication.

[0005] The A-type crystal may be prepared from a fumarate salt of compound (I) through recrystallization from a mixed solvent of methanol and isopropanol. The molar ratio between compound (I) and fumarate is 2:1. The A-type crystal provides the diffraction pattern as shown in Fig. 1 when measured by X-ray diffractometry with Cu-K $\alpha$  radiation.

[0006] The C-type crystal may be prepared from a fumarate salt of compound (I) through treatment with ethyl acetate. The molar ratio between compound (I) and fumarate is 1:1. The C-type crystal provides the diffraction pattern as shown in Fig. 2 when measured by X-ray diffractometry with Cu-K $\alpha$  radiation.

[0007] The D-type crystal may be prepared from a fumarate salt of compound (I) through treatment with a mixed solvent of ethyl acetate and water. The molar ratio between compound (I) and fumarate is 2:1. The D-type crystal provides the diffraction pattern as shown in Fig. 3 when measured by X-ray diffractometry with Cu-K $\alpha$  radiation.

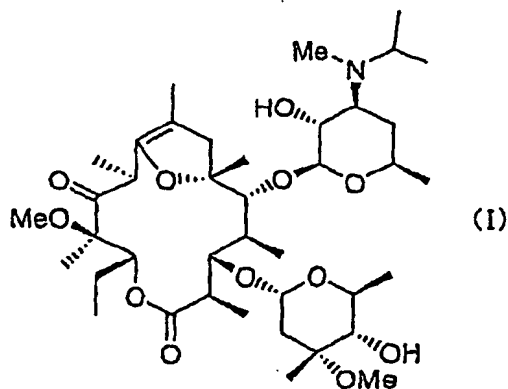
[0008] Among the A-type, C-type and D-type crystals, the D-type crystal is reported to have high quality as a pharmaceutical and a starting material therefor because it is superior in stability or other properties to the other crystals (JP 9-100291 A).

[0009] However, the prior art D-type crystal prepared by the conventionally known techniques as mentioned above which involve the following problems: a large volume of crystallization solvent remains in the crystal as a residual solvent; the residual solvent is difficult to remove during drying procedure; and the dryness of residual solvent cannot be below 1500 ppm. In this case, the residual solvent remaining in the D-type crystal is ethyl acetate, which is less toxic and less risky for human health (see "Guideline for residual solvents in pharmaceuticals" attached to the Notification No. 307 of March 30, 1998 delivered from the director of Evaluation and Licensing Division, Pharmaceutical and Medical Safety Bureau, Ministry of Health and Welfare, Japan). However, it is naturally more desirable to reduce the content of such a less toxic solvent remaining in the crystal in a case where the crystal is intended to be used as a starting material for pharmaceuticals. Preferably, the content of residual solvent should be reduced to 1500 ppm or below, more preferably 1000 ppm or below. The prior art D-type crystal also involves another problem of having a small particle size, which often leads to tableting troubles during the preparation of tablets comprising this crystal.

DISCLOSURE OF THE INVENTION

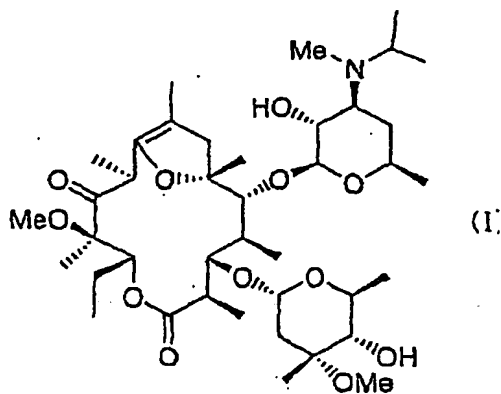
[0010] As a result of extensive and intensive efforts made to overcome the above problems in the prior art, the present inventors found structurally novel crystals of a fumarate salt of compound (I), which were different from known crystals, and finally completed the invention based on this finding.

[0011] Namely, the present invention is directed to an E-type crystal of a fumarate salt of the compound represented by formula (I):



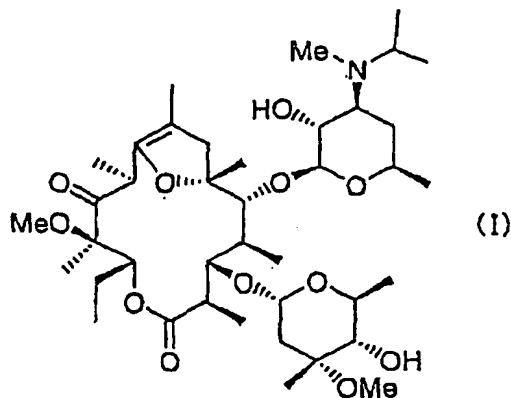
wherein said E-type crystal has strong X-ray diffraction peaks at diffraction angles ( $2\theta$ ) of  $5.6^\circ$  and  $10.4^\circ$  as measured by X-ray diffractometry with Cu-K $\alpha$  radiation.

[0012] The present invention is also directed to a method for preparing an E-type crystal of a fumarate salt of the compound represented by formula (I):



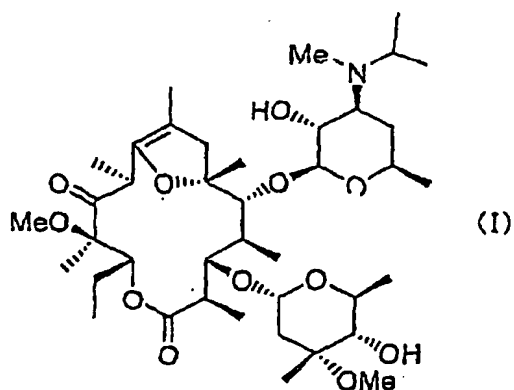
which comprises treating a C-type crystal of a fumarate salt of compound (I) in a mixed solvent of ethyl acetate and water at  $20^\circ\text{C}$  to  $40^\circ\text{C}$ , wherein said E-type crystal has strong X-ray diffraction peaks at diffraction angles ( $2\theta$ ) of  $5.6^\circ$  and  $10.4^\circ$  as measured by X-ray diffractometry with Cu-K $\alpha$  radiation.

[0013] Further, the present invention is directed to a D-type crystal of a fumarate salt of the compound represented by formula (I):



which is obtainable via an E-type crystal of a fumarate salt of compound (I) having strong X-ray diffraction peaks at diffraction angles ( $2\theta$ ) of  $5.6^\circ$  and  $10.4^\circ$  as measured by X-ray diffractometry with Cu-K $\alpha$  radiation.

[0014] Furthermore, the present invention is directed to a method for preparing a D-type crystal of a fumarate salt of the compound represented by formula (I):



which comprises obtaining the D-type crystal via an E-type crystal of a fumarate salt of compound (I) having strong X-ray diffraction peaks at diffraction angles ( $2\theta$ ) of  $5.6^\circ$  and  $10.4^\circ$  as measured by X-ray diffractometry with Cu-K $\alpha$  radiation.

#### BRIEF DESCRIPTION OF DRAWINGS

[0015]

Fig. 1 shows a powder X-ray diffraction pattern of the A-form crystal.

Fig. 2 shows a powder X-ray diffraction pattern of the C-form crystal.

Fig. 3 shows a powder X-ray diffraction pattern of the D-form crystal.

Fig. 4 shows a powder X-ray diffraction pattern of the E-form crystal.

Fig. 5 shows a DSC spectrum of the E-form crystal.

#### BEST MODE FOR CARRYING OUT THE INVENTION

[0016] The structurally novel crystal of a fumarate salt of the erythromycin derivative represented by formula (I) according to the present invention (hereinafter simply referred to as "E-type crystal") provides the diffraction pattern as shown in Fig. 4 when measured by X-ray diffractometry with Cu-K $\alpha$  radiation. Fig. 4 shows strong peaks at diffraction angles ( $2\theta$ ) of  $5.6^\circ$  and  $10.4^\circ$ .

[0017] These X-ray diffraction angles can be measured using various devices that are commercially available, such as a powder X-ray diffractometer. The details on the measurement principle of powder X-ray diffractometry can be found in the Japanese Pharmacopoeia, 13th Edition, published by Hirokawa Publishing Co. (1996) pp. B471-B475, the Japanese Pharmacopoeia, 14th Edition, published by Hirokawa Publishing Co. (2001) pp. B614-B619 and elsewhere. In general, the diffraction angle has an acceptable error of around  $\pm 0.2^\circ$ .

[0018] As used herein, the term "dryness" refers to the content of residual solvent which reaches an almost unchanged level during drying procedure, more specifically refers to the content at a point where the drying procedure produces a less than 100 ppm/hr decrease in the content of residual solvent.

[0019] The present invention will be described below in more detail.

[0020] The method for preparing the E-type crystal according to the present invention can start with the C-type crystal.

[0021] The C-type crystal may be prepared by treating the A-type crystal with ethyl acetate, as described in JP 9-100241 A.

[0022] Further, the A-type crystal may be prepared by treating a fumarate salt of compound (I) with a mixed solvent of methanol and isopropanol, as described in JP 6-56873 A and JP 9-100241 A.

[0023] The E-type crystal of the present invention can be prepared by suspending the C-type crystal in a mixed solvent of ethyl acetate and water. The C-type crystal may be used either in isolated crystal or as a suspension in the solvent, but preferably used as a suspension in the solvent. In a preferred embodiment, the A-type crystal is treated with ethyl acetate to yield the C-type crystal as a suspension in ethyl acetate, which is further suspended by addition of water.

[0024] In the mixed solvent used in this suspension procedure, the mixing ratio between ethyl acetate and water is usually set to 99:1 to 95:5, preferably 97:3 to 95:5. The suspension procedure is usually performed at a temperature of  $20^\circ\text{C}$  to  $40^\circ\text{C}$ , preferably  $20^\circ\text{C}$  to  $30^\circ\text{C}$ . A temperature below  $20^\circ\text{C}$  tends to stimulate the conversion into the D-type crystal. The suspension procedure is usually continued for 30 minutes to 300 minutes, preferably 60 minutes to 240 minutes.

[0025] The resulting E-type crystal may be separated from the solvent by filtration, centrifugation or the like. The separated E-type crystal may be dried under reduced pressure or other conditions, but preferably dried under reduced pressure. The drying temperature is usually  $20^\circ\text{C}$  to  $60^\circ\text{C}$ , preferably  $30^\circ\text{C}$  to  $50^\circ\text{C}$ .

[0026] The E-type crystal may be suspended in a mixed solvent of ethyl acetate and water at a temperature below  $20^\circ\text{C}$  to yield the D-type crystal. In the mixed solvent used here, the mixing ratio between ethyl acetate and water is preferably set to 99:1 to 97:3. The suspension procedure is preferably performed at a temperature of  $-20^\circ\text{C}$  to  $20^\circ\text{C}$  and usually continued for 1 hour to 12 hours, preferably 3 hours to 11 hours, more preferably 5 hours to 10 hours.

[0027] In order to prepare the D-type crystal with an average particle size sufficient to avoid tableting troubles (preferably  $90\text{ }\mu\text{m}$  or more, more preferably  $100\text{ }\mu\text{m}$  or more) from the E-type crystal, the mixing ratio between ethyl acetate and water is preferably set to 98.1:1.9 to 97:3 in the mixed solvent used for crystallization. The suspension procedure is performed at a temperature of  $10^\circ\text{C}$  to  $20^\circ\text{C}$ , preferably  $11^\circ\text{C}$  to  $19^\circ\text{C}$ , more preferably  $13^\circ\text{C}$  to  $18^\circ\text{C}$ . In order to stimulate the conversion into the D-type crystal or to improve the yield of the D-type crystal, the suspension may further be cooled to  $-20^\circ\text{C}$  to  $10^\circ\text{C}$ , preferably  $-15^\circ\text{C}$  to  $10^\circ\text{C}$ . The suspension procedure is usually continued for several minutes to 20 hours, preferably 5 minutes to 4 hours, more preferably 10 minutes to 2 hours. In a case where the suspension is further cooled, the suspension procedure is usually continued for additional several minutes to 20 hours, preferably around 1 hour.

[0028] It should be understood that the period of time required for the individual suspension procedures mentioned above refers to the minimum period of time required to prepare the E-type crystal, required to prepare the D-type crystal from the E-type crystal and required to prepare the D-type crystal with a large average particle size from the E-type crystal. The individual suspension procedures may be continued beyond the minimum period of time, depending on the degree of crystal growth or the convenience of preparation steps.

[0029] In preparing the D-type crystal via the E-type crystal, the D-type crystal can also be prepared continuously from the C-type crystal via the E-type crystal by merely controlling the temperature, without isolating the E-type crystal during preparation.

[0030] The resulting D-type crystal may be separated from the solvent by filtration, centrifugation or the like, and then dried under reduced pressure. The drying temperature is preferably  $20^\circ\text{C}$  to  $70^\circ\text{C}$ . The D-type crystal of the present invention may completely (100%) or partially be composed of compound molecules prepared via the E-type crystal. In the latter case, the D-type crystal prepared via the E-type crystal may be contained at any percentage as long as the content of residual solvent does not exceed 1500 ppm, preferably not exceed 1000 ppm, and/or tableting troubles do not occur.

[0031] The D-type crystal prepared via the E-type crystal in this way ensures a residual solvent content of 1500 ppm or below, which could not be achieved by the prior art D-type crystal. In addition, the D-type crystal thus prepared further ensures a residual solvent content of 1000 ppm or below and is also easier to dry than the prior art D-type crystal; it is therefore more preferable as an active pharmaceutical ingredient. The D-type crystal partially prepared via

E-type crystal also ensures a residual solvent content of 1500 ppm or below, which could not be achieved by the prior art D-type crystal. This D-type crystal further ensures a residual solvent content of 1000 ppm or below and is also easier to dry than the prior art D-type crystal; it is also therefore more preferable as an active pharmaceutical ingredient.

[0032] Further, the D-type crystal with a large average particle size prepared under the conditions mentioned above cannot be obtained by the prior art techniques and allows avoidance of tableting troubles; it is therefore particularly advantageous in preparing pharmaceuticals.

[0033] The content of residual solvent may be determined in a known manner, for example, by gas chromatography. The details on gas chromatography can be found in the Japanese Pharmacopoeia, 13th Edition, published by Hirokawa Publishing Co. (1996) pp. B83-B94, the Japanese Pharmacopoeia, 14th Edition, published by Hirokawa Publishing Co. (2001) pp. B98-B114 and elsewhere. In general, gas chromatography will cause a measurement error falling within around  $\pm 1\%$ . The compound of the present invention can also be determined for its average particle size in a known manner or using various devices that are commercially available, such as a dry particle size distribution analyzer. In general, such a particle size distribution analyzer will cause a measurement error falling within around  $\pm 5\%$ .

## EXAMPLES

[0034] The present invention will be further described in the following examples, which are provided for illustrative purposes only and are not intended to limit the scope of the invention.

[0035] In these examples and comparison examples, X-ray diffractometry was performed using a powder X-ray diffractometer RINT-1100 (Rigaku), the content of residual solvent was determined using a gas chromatograph GC-17A (Shimadzu Corp.), and the average particle size was determined using a particle size distribution analyzer RPS-95 (Seishin Enterprise Co., Ltd.). The content of residual solvent had a measurement error of around  $\pm 1\%$ , while the average particle size had a measurement error of around  $\pm 4\%$ .

### [Example 1]

[0036] A fumarate salt of compound (I) (50.0 g) was dissolved in ethyl acetate (400 mL) and methanol (40 mL) at room temperature. The solution was then concentrated to dryness under reduced pressure at room temperature. The resulting dried product was stirred in ethyl acetate (415 mL) at 25°C for 1 hour to give a suspension of a C-form crystal. Water (4.15 mL) was added to this suspension, followed by stirring at 25°C for 0.5 hours. Water (4.15 mL) was further added and stirred at 25°C for 0.5 hours. Water (4.15 mL) was further added and stirred at 25°C for 0.5 hours. The suspension was then cooled to 20°C, stirred for 1 hour, and filtered to give a wet crystal (43.7 g). This wet crystal was dried under reduced pressure at 40°C for 3 hours to give a crystal of the fumarate salt of compound (I) (34.1 g). This crystal was confirmed to be an E-form crystal having strong peaks at diffraction angles (2 $\theta$ ) of 5.6° and 10.4° as measured by X-ray diffractometry. Fig. 5 shows a DSC spectrum of the resulting E-form crystal.

### [Example 2]

[0037] A fumarate salt of compound (I) (20 g) was stirred in ethyl acetate (166 mL) at 25°C for 2 hours to give a C-form crystal. After addition of water (2.4%, 4.0 mL), this C-form crystal was gradually cooled to ensure its complete conversion into E-form. The suspension was further cooled to 15°C and stirred for 3 hours, followed by cooling to -10°C. The resulting crystal was then isolated to give a wet D-form crystal (20.4 g, average particle size: 302  $\mu$ m). This wet D-form crystal was dried under reduced pressure at 25°C for 1 hour and further dried at 60°C to give the D-form crystal of the fumarate salt of compound (I). The resulting D-form crystal was found to have a residual solvent content of 78 ppm.

### [Example 3]

[0038] Starting with a C-form crystal of a fumarate salt of compound (I), the same procedures as shown in Example 2 were repeated, except for adding 2.6% water, to give a D-form crystal with a particle size of 197  $\mu$ m via an E-form crystal.

### [Example 4]

[0039] A fumarate salt of compound (I) (11.6 kg) was dissolved at 25°C in a mixed solvent of ethyl acetate (104.6 kg) and methanol (9.2 kg). After the solution was concentrated, ethyl acetate (86.8 kg) was added at 25°C to the concentrated residue, followed by stirring at 24°C for 1 hour to give a C-form crystal. After addition of water (2.0%, 1.9

kg), this C-form crystal was gradually cooled to ensure its conversion into E-form crystal. The suspension was further cooled to 15°C and stirred for 1 hour, followed by cooling to -10°C. The resulting crystal was then centrifuged to give a wet D-form crystal (13.4 kg). This wet D-form crystal was dried under reduced pressure at 60°C for 28 hours to give the D-form crystal of the fumarate salt of compound (I) (10.5 kg, yield 90.5%, average particle size: 141 µm). The resulting D-form crystal was found to have a residual solvent content of 988 ppm. In addition, no tableting trouble was observed in this D-form crystal when used as a main component to prepare tablets.

[Example 5]

**[0040]** A fumarate salt of compound (I) (11.6 kg) was dissolved at 30°C in a mixed solvent of ethyl acetate (94.2 kg) and methanol (9.1 kg). After the solution was concentrated, ethyl acetate (86.8 kg) was added at 22°C to the concentrated residue, followed by stirring at 24°C for 1 hour to give a C-form crystal. After addition of water (2.0%, 1.9 kg), this C-form crystal was gradually cooled to ensure its conversion into E-form crystal. The suspension was further cooled to 15°C and stirred for 1 hour, followed by cooling to -10°C. The resulting crystal was then centrifuged to give a wet D-form crystal (13.2 kg). This wet D-form crystal was dried under reduced pressure at 60°C for 10 hours to give the D-form crystal of the fumarate salt of compound (I) (10.5 kg, yield 90.5%, average particle size: 197 µm). The resulting D-form crystal was found to have a residual solvent content of 845 ppm. In addition, no tableting trouble was observed in this D-form crystal when used as a main component to prepare tablets.

**[0041]** C-form crystals of a fumarate salt of compound (I) were similarly treated in accordance with Example 4 or 5 to give D-form crystals via E-form crystals (Examples 6 to 8).

[Comparison Example 1]

**[0042]** A fumarate salt of compound (I) (10.8 kg) was dissolved at 25°C in a mixed solvent of ethyl acetate (87.6 kg) and methanol (8.5 kg). After the solution was concentrated, ethyl acetate (80.8 kg) was added at 25°C to the concentrated residue, followed by stirring at 25°C for 1 hour to give a C-form crystal. After addition of water (1.5%, 1.3 kg), this C-form crystal was cooled to 15°C and stirred for 1 hour to ensure its conversion into D-form crystal. The suspension was further cooled to -10°C and stirred for 1 hour. The resulting crystal was then centrifuged to give a wet D-form crystal (12.7 kg). This wet D-form crystal was dried under reduced pressure at 60°C for 16 hours to give the D-form crystal of the fumarate salt of compound (I) (10.8 kg, yield 87.4%, average particle size: 82 µm). Tableting troubles were observed in this D-form crystal when used as a main component to prepare tablets.

**[0043]** A C-form crystal of a fumarate salt of compound (I) was similarly treated in accordance with Comparison Example 1 to give a D-form crystal without going via an E-form crystal (Comparison Example 2).

**[0044]** Table 1 summarizes the properties of the D-type crystals prepared via E-type crystals (Examples 2 to 8) and the prior art D-type crystals (Comparison Examples 1 and 2).

Table 1

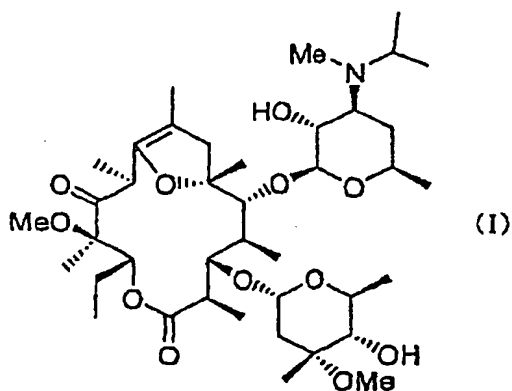
	Percentage of water (%)	Prepared via E-type crystal	Conditions for conversion of E-type crystal into D-type crystal	Drying conditions	Content of residual solvent (ppm)	Particle size ( $\mu\text{m}$ )	Tabletting trouble
Example 2	2.4	Completely	15°C, 3hr $\rightarrow$ -10°C	Reduced pressure 60°C, 8hr	78	302	---
Example 3	2.6	Completely	15°C, 6hr $\rightarrow$ -10°C	Reduced pressure 60°C, 8hr	---	197	---
Example 4	2.0	Partially	15°C, 1hr $\rightarrow$ -10°C	Reduced pressure 60°C, 28hr	988	141	No
Example 5	2.0	Partially	15°C, 1hr $\rightarrow$ -10°C	Reduced pressure 60°C, 10hr	845	197	No
Example 6	2.0	Partially	13°C, 0.5hr $\rightarrow$ -10°C	Reduced pressure 60°C, 9hr	1049	---	---
Example 7	2.0	Partially	15°C, 1hr $\rightarrow$ -10°C	Reduced pressure 60°C, 6hr	647	163	No
Example 8	2.0	Partially	15°C, 1hr $\rightarrow$ -10°C	Reduced pressure 60°C, 10hr	893	185	No
Comparison Example 1	1.5	Not	---	Reduced pressure 60°C, 16hr	2228	82	Yes
Comparison Example 2	1.5	Not	---	Aeration 45°C, 20hr	1610	61	Yes

INDUSTRIAL APPLICABILITY

[0045] The E-form crystal of a fumarate salt of compound (I) according to the present invention enables the preparation of the D-form crystal with superior properties including a reduced content of residual solvent and high suitability for formulation. More specifically, the E-form crystal is characterized by (1) providing pharmaceuticals with superior quality and (2) allowing the efficient preparation of pharmaceuticals; it is therefore extremely useful in pharmaceutical preparation.

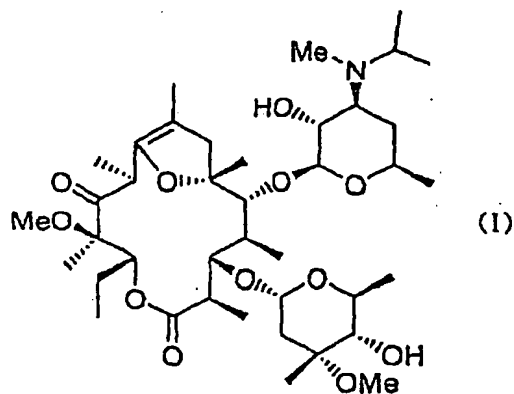
**Claims**

1. An E-type crystal of a fumarate salt of the compound represented by formula (I):



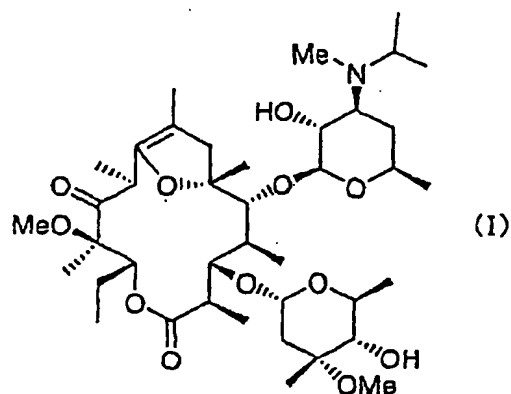
wherein said E-type crystal has strong X-ray diffraction peaks at diffraction angles ( $2\theta$ ) of  $5.6^\circ$  and  $10.4^\circ$  as measured by X-ray diffractometry with Cu-K $\alpha$  radiation.

2. A method for preparing an E-type crystal of a fumarate salt of the compound represented by formula (I):



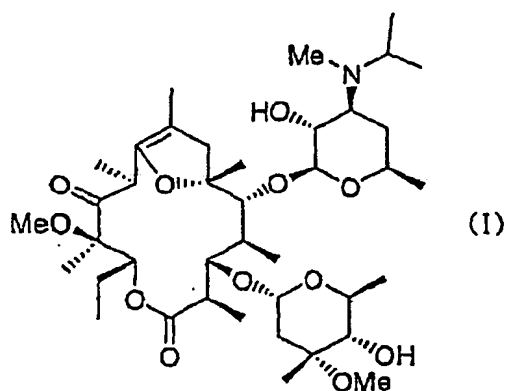
which comprises treating a C-type crystal of a fumarate salt of compound (I) in a mixed solvent of ethyl acetate and water at  $20^\circ\text{C}$  to  $40^\circ\text{C}$ , wherein said E-type crystal has strong X-ray diffraction peaks at diffraction angles ( $2\theta$ ) of  $5.6^\circ$  and  $10.4^\circ$  as measured by X-ray diffractometry with Cu-K $\alpha$  radiation.

3. A D-type crystal of a fumarate salt of the compound represented by formula (I):



20 which is obtainable via an E-type crystal of a fumarate salt of compound (I) having strong X-ray diffraction peaks at diffraction angles ( $2\theta$ ) of  $5.6^\circ$  and  $10.4^\circ$  as measured by X-ray diffractometry with Cu-K $\alpha$  radiation.

- 25
- 30
4. The D-type crystal according to claim 3, wherein the crystal has a residual solvent content of 1500 ppm or below.
  5. The D-type crystal according to claim 3, wherein the crystal has a residual solvent content of 1000 ppm or below.
  6. The compound according to any one of claims 3 to 5, wherein the crystal has an average particle size of 90  $\mu\text{m}$  or more.
  7. The compound according to any one of claims 3 to 5, wherein the crystal has an average particle size of 100  $\mu\text{m}$  or more.
  8. A method for preparing a D-type crystal of a fumarate salt of the compound represented by formula (I):



50 which comprises obtaining the D-type crystal via an E-type crystal of a fumarate salt of compound (I) having strong X-ray diffraction peaks at diffraction angles ( $2\theta$ ) of  $5.6^\circ$  and  $10.4^\circ$  as measured by X-ray diffractometry with Cu-K $\alpha$  radiation.

- 55
9. The method according to claim 8, wherein the D-type crystal of a fumarate salt of compound (I) has a residual solvent content of 1500 ppm or below.
  10. The method according to claim 8, wherein the D-type crystal of a fumarate salt of compound (I) has a residual

solvent content of 1000 ppm or below.

- 5
11. The method according to any one of claims 8 to 10, wherein the D-type crystal of a fumarate salt of compound (I) has an average particle size of 90  $\mu\text{m}$  or more.
12. The method according to any one of claims 8 to 10, wherein the D-type crystal of a fumarate salt of compound (I) has an average particle size of 100  $\mu\text{m}$  or more.
- 10
- 15
- 20
- 25
- 30
- 35
- 40
- 45
- 50
- 55

Fig. 1

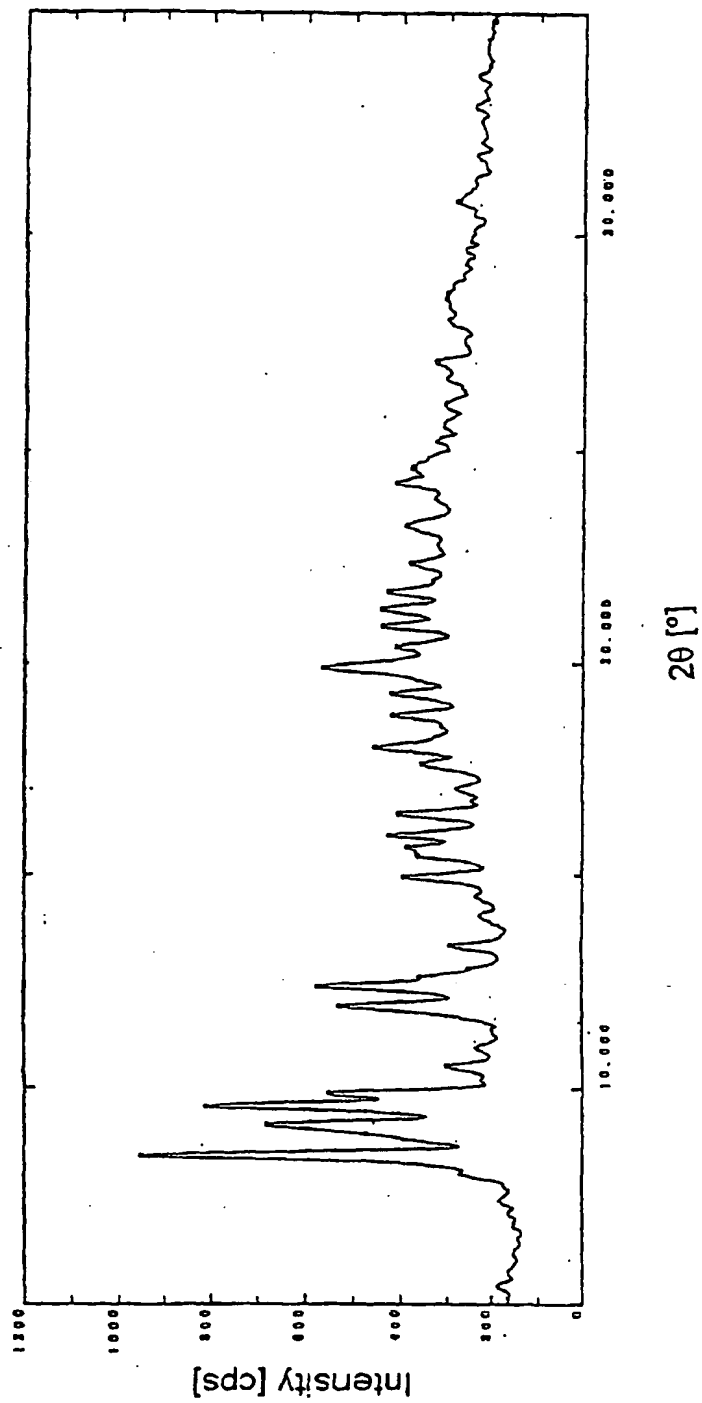


Fig. 2

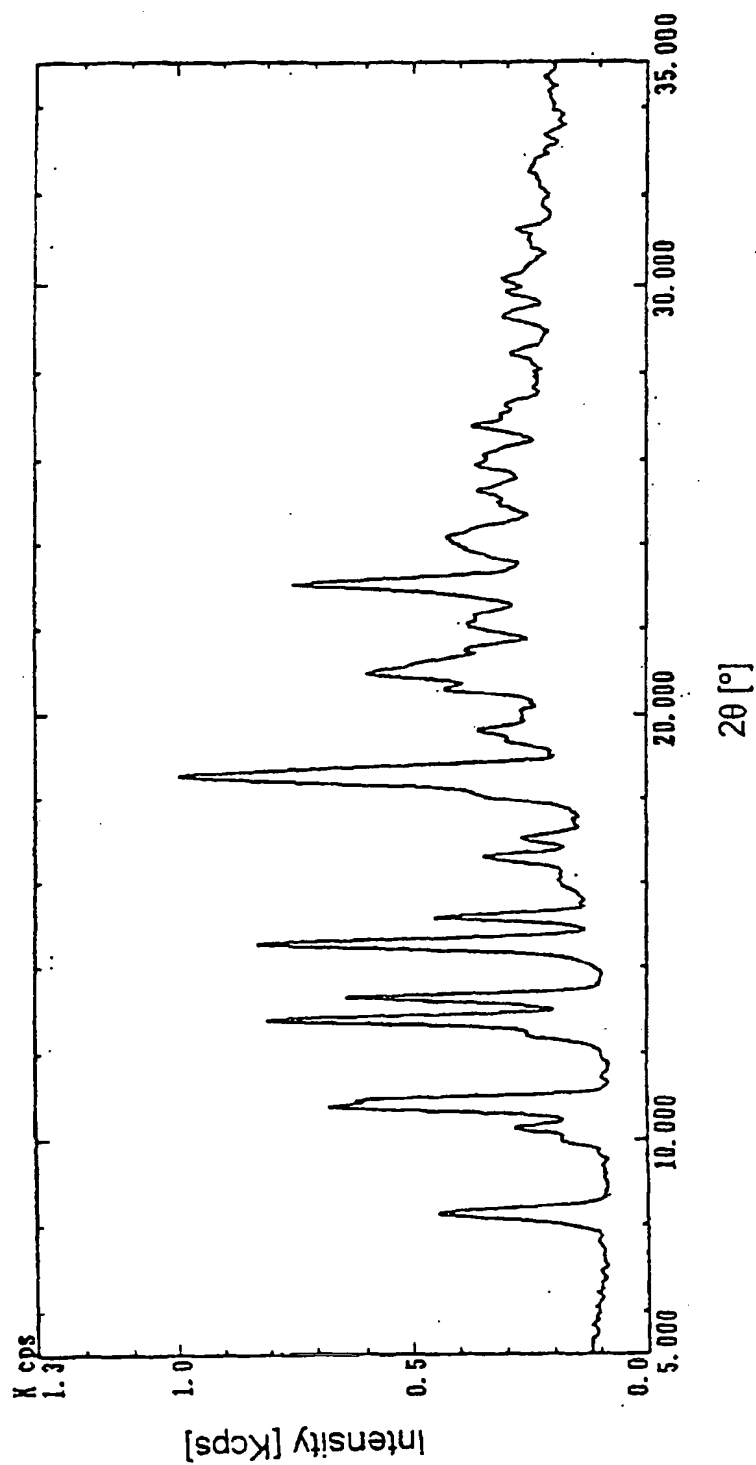


Fig. 3

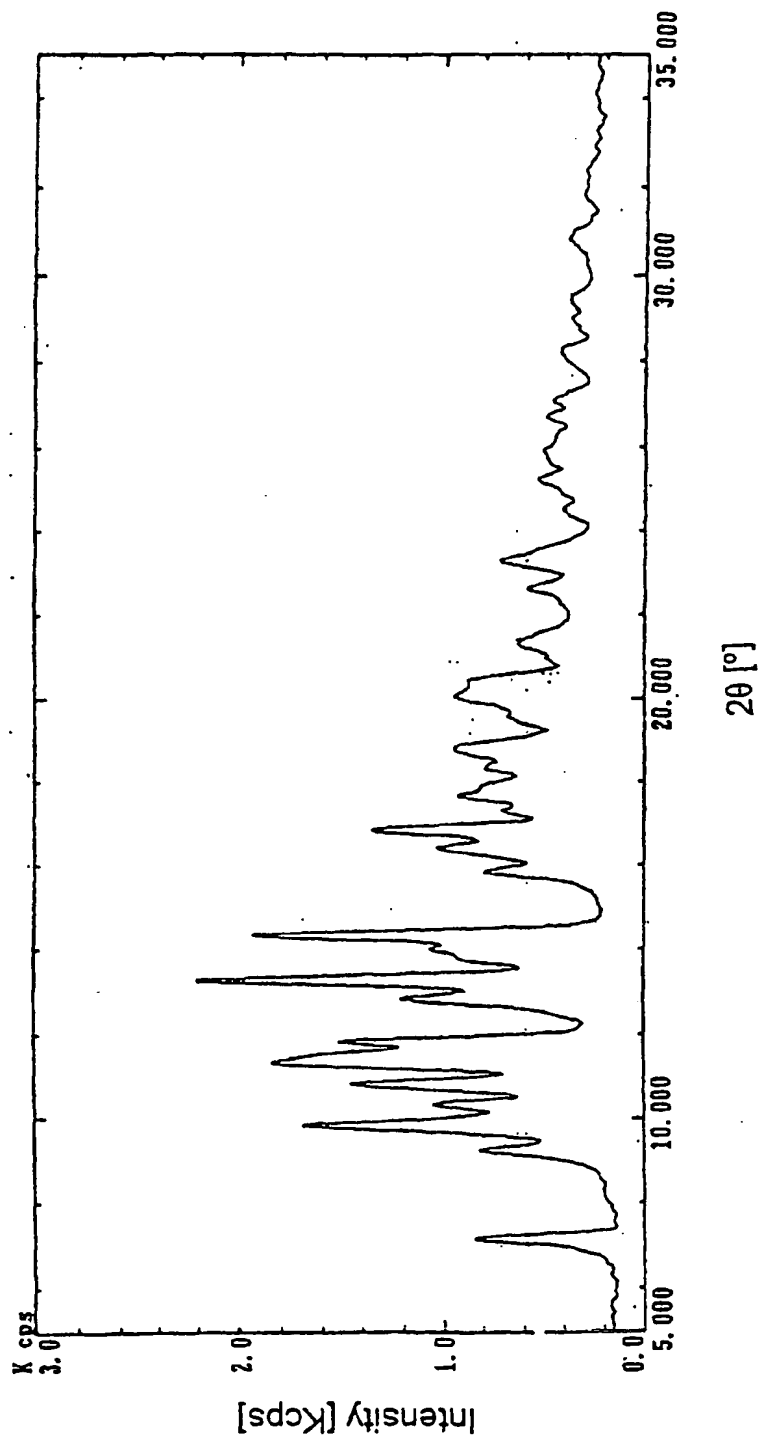


Fig. 4

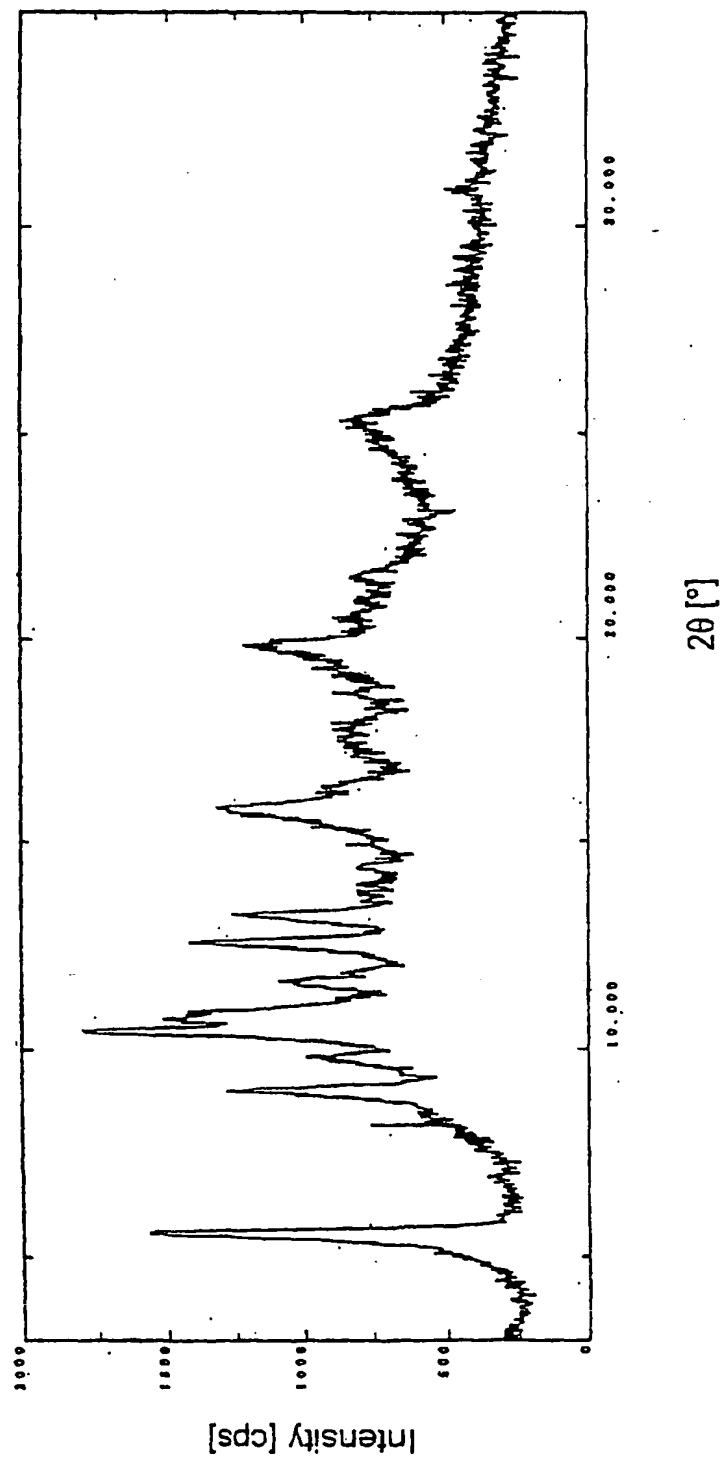
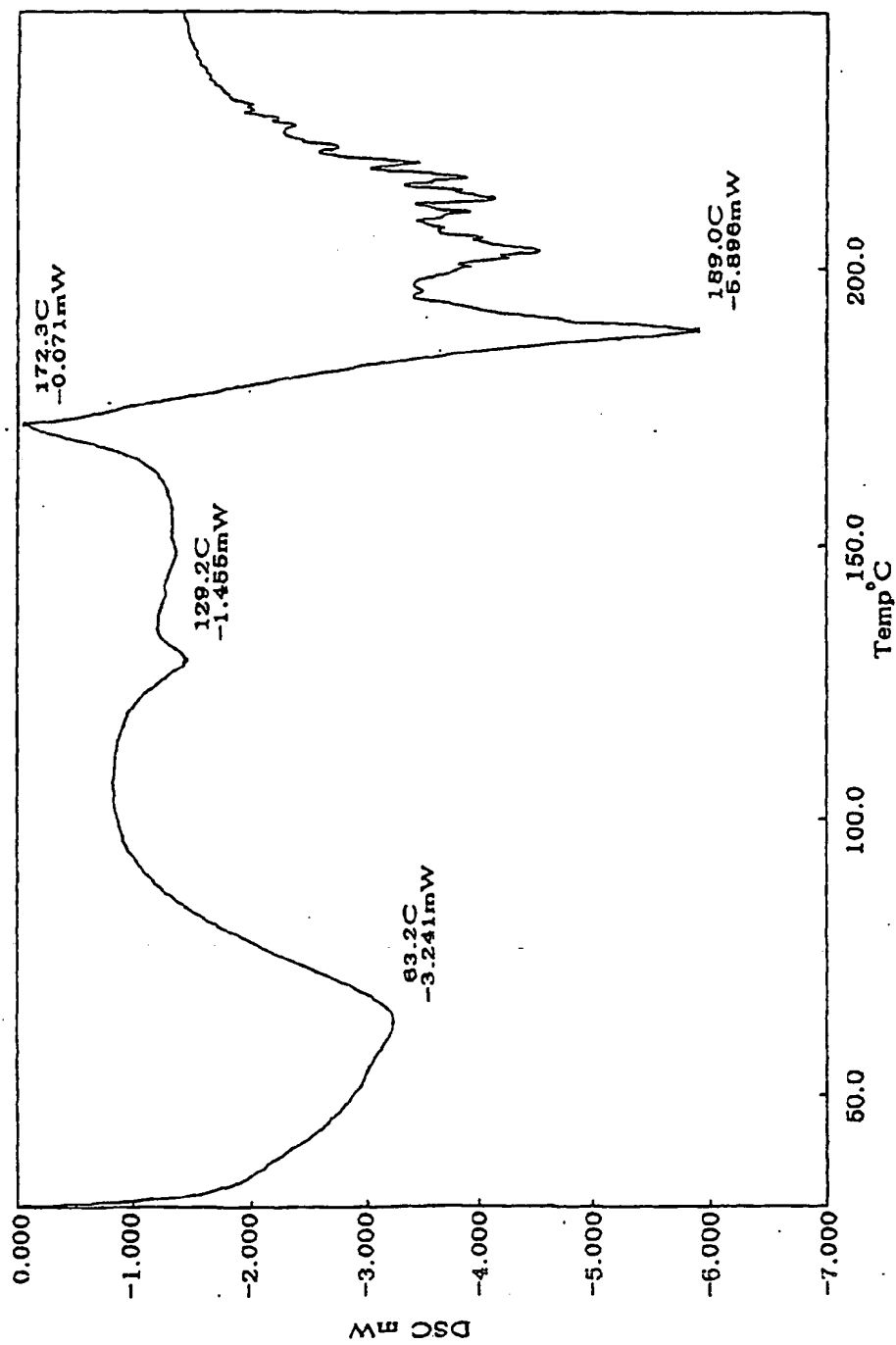


Fig. 5



## INTERNATIONAL SEARCH REPORT

International application No.

PCT/JP01/08990

<b>A. CLASSIFICATION OF SUBJECT MATTER</b> Int.Cl. <sup>7</sup> C07H17/08 // A61K31/7048, A61P1/00														
According to International Patent Classification (IPC) or to both national classification and IPC														
<b>B. FIELDS SEARCHED</b> Minimum documentation searched (classification system followed by classification symbols) Int.Cl. <sup>7</sup> C07H17/08														
Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched														
Electronic data base consulted during the international search (name of data base and, where practicable, search terms used) REGISTRY (STN) , CAPLUS (STN) , CAOLD (STN)														
<b>C. DOCUMENTS CONSIDERED TO BE RELEVANT</b> <table border="1"> <thead> <tr> <th>Category*</th> <th>Citation of document, with indication, where appropriate, of the relevant passages</th> <th>Relevant to claim No.</th> </tr> </thead> <tbody> <tr> <td>X</td> <td>EP 846697 A1 (SHUGAI SEIYAKU KABUSHIKI KAISHA) , 10 June, 1998 (10.06.98) , Full text</td> <td>3-7</td> </tr> <tr> <td>A</td> <td>&amp; JP 9-100291 A                      &amp; WO 97/06177 A1 &amp; IL 119002 A1                      &amp; CA 2228254 AA &amp; AU 9666306 A1                      &amp; CN 1196058 A &amp; US 5959088 A</td> <td>1, 2, 8-12</td> </tr> <tr> <td>A</td> <td>EP 643068 A1 (SHUGAI SEIYAKU KABUSHIKI KAISHA) , 15 March, 1995 (15.03.95) , &amp; JP 6-56873 A                      &amp; WO 93/24509 A1 &amp; IL 105810 A1                      &amp; ZA 9303679 A &amp; AU 9340897 A1                      &amp; CN 1081184 A &amp; US 5658888 A                      &amp; AT 169927 E &amp; ES 2120499 T3</td> <td>1-12</td> </tr> </tbody> </table>			Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.	X	EP 846697 A1 (SHUGAI SEIYAKU KABUSHIKI KAISHA) , 10 June, 1998 (10.06.98) , Full text	3-7	A	& JP 9-100291 A                      & WO 97/06177 A1 & IL 119002 A1                      & CA 2228254 AA & AU 9666306 A1                      & CN 1196058 A & US 5959088 A	1, 2, 8-12	A	EP 643068 A1 (SHUGAI SEIYAKU KABUSHIKI KAISHA) , 15 March, 1995 (15.03.95) , & JP 6-56873 A                      & WO 93/24509 A1 & IL 105810 A1                      & ZA 9303679 A & AU 9340897 A1                      & CN 1081184 A & US 5658888 A                      & AT 169927 E & ES 2120499 T3	1-12
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